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Views Themes Grouping & Totals Controls Header / Footer Tools

All Access Objects

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- qryManagementPlan

qryCurrentProjects

Saturday, July 2, 2016 1:09:30 AM

Project Name	Project Status	ProjectStart	ProjectEnd	Number of Late Tasks	OnTime
Nerves Too Late: Reconnecting with Your Adult Children	In Progress	1/26/2013			On Time
The Petals, the Seeds, and the Cradles	In Progress	1/26/2013		20	Late
The Great American Frontier	Waiting on Approval				On Time
Great American Beaches	Waiting on Approval				On Time
Cash is King: How to Cut Your Spending by Curbing Cash	In Progress	6/10/2013		2	Late
Greatest Bandwagons of the 20th Century	On Hold	6/25/2012			On Time
The Snake in the Grass	Waiting on Approval				On Time
The Light of Heart	Not Started				On Time
Haunt of Someone	On Hold	2/25/2013			On Time
Growing Up Nobody	In Progress	3/29/2013			On Time
Willow of Dreams	In Progress	2/26/2013			On Time
Visions of Danger	On Hold	4/29/2013			On Time
The River in the Throat	On Hold	3/2/2013			On Time
The Soul in the Crying	Not Started				On Time
The Memory in the Man	Not Started				On Time

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web donor whereas a base generically represented by b is a proton acceptor when acids and bases combine they swap a proton $ha + b \rightleftharpoons a^- + hb$ strong acids e.g. sulfuric acid have a very pronounced tendency to transfer a proton to a base whereas weak acids

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web understand that parts of an acid base titration be able to determine the K_a or K_b from pH data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate K_{a1} and K_{a2} for a polyprotic acid by the end of this lab students should be able to

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web answer based on the data the strong acid strong base combination of hydrochloric acid and sodium hydroxide had the neutral pH at the equivalence point being 7 this result is expected because a combination of strong acids and strong bases should result in a neutral at the equivalence point

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web 1 approximately 3 g of hydrated oxalic acid $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$ was weighted to the nearest 0.01 g in a 50 ml beaker 2 about 30 ml of distilled water was added to the beaker to dissolve the oxalic acid 3 the solution was transferred into 250 ml volumetric flask the beaker was rinsed and its contents were poured into the flask

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