

## 17.1

## HYDROGEN IONS AND ACIDITY

## Section Overview

## Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and an appropriate equilibrium.
2. Calculate hydroxide ion concentration, hydronium ion concentration, and pH of an acid or base solution.
3. Calculate the pH of weak acid solutions.

## Vocabulary

- |                   |                             |                   |
|-------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid      |
| 4. autoionization | 5. hydronium ion            | 6. autoionization |
|                   | 7. weak acid                | 8. pH             |

## Key Equations

- |  |  |
|--|--|
| 1. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ | 2. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ |
| 3. $\text{pH} = -\log[\text{H}_3\text{O}^+]$   | 4. $\text{pH} = -\log[\text{H}_3\text{O}^+]$   |

## Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in the second step. Each choice is used only once. Complete the table.

- |  |       |
|--|-------|
| 1. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (water autoionization)  | _____ |
| 2. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (the autoionization of water which produces an acidic solution) | _____ |
| 3. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (autoionization of water)                                       | _____ |
| 4. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (the autoionization of water)                                   | _____ |
| 5. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (the autoionization of water)                                   | _____ |
| 6. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (the autoionization of water)                                   | _____ |
| 7. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (the autoionization of water)                                   | _____ |
| 8. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (the autoionization of water)                                   | _____ |

- |       |                             |
|-------|-----------------------------|
| _____ | 1. autoionization           |
| _____ | 2. autoionization of water  |
| _____ | 3. autoionization of water  |
| _____ | 4. autoionization of water  |
| _____ | 5. autoionization of water  |
| _____ | 6. autoionization of water  |
| _____ | 7. autoionization of water  |
| _____ | 8. autoionization of water  |
| _____ | 9. autoionization of water  |
| _____ | 10. autoionization of water |
| _____ | 11. autoionization of water |
| _____ | 12. autoionization of water |
| _____ | 13. autoionization of water |
| _____ | 14. autoionization of water |

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**Ensheng Dong**



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