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which is equal to 22.4 l for 1 mole of any ideal gas at a temperature equal to 273.15 k and a pressure

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web the molar volume will be the volume occupied by one molar mass 32 g of the liquid expressing the density in liters  
instead of ml we have  $v = m \cdot d$   $\frac{32 \text{ g}}{\text{mol}} = 1.790 \text{ g/l} = 0.0405 \text{ l/mol}$  nonnumber

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point to keep in mind is that n can be moles of any gas since one mole always contain 6.02  $\times 10^{23}$  molecules and one molecule  
of any gas at the same temperature will have the same kinetic energy based on this principle it is found

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number an unchanging constant whether you are counting hydrogen atoms carbon dioxide molecules or carbonate ions a  
mole is always 6.02  $\times 10^{23}$  but as different particles have different mass their molar mass varies

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web the molar volume has the si unit of cubic metres per mole  $\text{m}^3 \text{ mol}^{-1}$  although it is more typical to use the units cubic decimetres per mole  $\text{dm}^3 \text{ mol}^{-1}$  for gases and cubic centimetres per mole  $\text{cm}^3 \text{ mol}^{-1}$  for liquids and solids

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web zero degrees celsius or 273 k is defined as standard temperature and one atmosphere or 760 mm hg is defined as standard pressure the volume occupied by one mole of a gas is called molar volume therefore the molar volume of a gas at standard temperature and pressure stp is 22.4 l

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web at standard temperature and pressure stp the molar volume  $V_m$  is the volume occupied by one mole of a chemical element or a chemical compound it can be calculated by dividing the molar mass  $M$  by mass density  $\rho$  molar gas volume is one mole of any gas at a specific temperature and pressure has a fixed volume molar volume formula

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