

Chemistry 2nd Semester Final Review **KEY**

This is meant as a guide. It does NOT contain every single thing you need to know for the final exam. Review notes, benchmark review sheets, homeworks, chapter study guides, etc.

Bring your book to school by Monday for turn in (seniors should bring book to bookstore by themselves ASAP)

Gas Laws (Ch. 13):

Subtopics: atmospheric pressure, Dalton's law of partial pressures, combined gas law, Ideal gas law, gas stoichiometry (22.4 L/mol for a gas at STP)

Practice questions:

- If a gas occupies 3.8 L at a pressure of 2.71 atm, what would the pressure be if the volume changes to 1.47 L? (Temp. and amt. of gas constant)
 $V_1=3.8 \text{ L}; P_1=2.71 \text{ atm}$
 $V_2=1.47 \text{ L}; P_2=? \text{ Atm}$
Since all units are standard (L, atm), we do not need to convert the units before proceeding.
 $(P_1V_1)/T_1=(P_2V_2)/T_2$
T goes away (not mentioned in problem)
 $P_1V_1=P_2V_2$
Isolate for P₂: $P_2=(P_1V_1)/V_2$
 $P_2=(2.71 \text{ atm} \times 3.8 \text{ L})/(1.47 \text{ L})=7.0 \text{ atm}$ (L cancelled and it's 2 sig figs)
- A gas has a volume of 1.49 L at a temperature of 34.75 degrees Celsius. What would the volume be at 78.41 degrees Celsius? (pressure & amt. of gas constant)
 $V_1=1.49 \text{ L}, T_1=34.75 \text{ degrees Celsius}$
 $V_2=? \text{ L}, T_2=78.41 \text{ degrees Celsius}$
We need to convert from Celsius to Kelvins before proceeding.
 $T_1=34.75 + 273=307.75 \text{ K}$
 $T_2=78.41 + 273=351.41 \text{ K}$
 $(P_1V_1)/T_1=(P_2V_2)/T_2$
Get rid of P
 $V_1/T_1=V_2/T_2$
Isolate V₂: $V_2=(V_1/T_1)T_2$
 $V_2=(1.49 \text{ L}/307.75 \text{ K})(351.41 \text{ K})=1.70 \text{ L}$
- What volume is occupied by 8.47 g of hydrogen gas at 84.7 degrees Celsius and 1.04 atm?
The combined gas law won't work here. You only have one condition and nothing changes.
 $PV=nRT$ will work. (n= moles and R= 0.08206 L atm/K mol)
 $P=1.04 \text{ atm}, V=? \text{ L}, n=?$ OOPS we need to convert from grams to moles!, $T=84.7 \text{ degrees Celsius} + 273=357.7 \text{ K}$
n: Hydrogen gas is diatomic!
 $8.47 \text{ g H}_2 \times (1 \text{ mol H}_2/2.02 \text{ g H}_2)=4.193069307 \text{ mol H}_2$
Isolate V: $V=(nRT)/P=(4.193069307 \text{ mol H}_2 \times 0.08206 \text{ L atm/K mol} \times 357.7 \text{ K})/1.04 \text{ atm}=118 \text{ L}$
- What volume is occupied by 56.75 g of oxygen gas at STP?

Nd Semester Final Review Guide Chemistry

Rosina Ehmann



Nd Semester Final Review Guide Chemistry:

The Top Books of the Year Nd Semester Final Review Guide Chemistry The year 2023 has witnessed a noteworthy surge in literary brilliance, with numerous compelling novels captivating the hearts of readers worldwide. Lets delve into the realm of top-selling books, exploring the engaging narratives that have charmed audiences this year. The Must-Read : Colleen Hoover's "It Ends with Us" This poignant tale of love, loss, and resilience has captivated readers with its raw and emotional exploration of domestic abuse. Hoover skillfully weaves a story of hope and healing, reminding us that even in the darkest of times, the human spirit can succeed. Nd Semester Final Review Guide Chemistry : Taylor Jenkins Reids "The Seven Husbands of Evelyn Hugo" This spellbinding historical fiction novel unravels the life of Evelyn Hugo, a Hollywood icon who defies expectations and societal norms to pursue her dreams. Reids absorbing storytelling and compelling characters transport readers to a bygone era, immersing them in a world of glamour, ambition, and self-discovery. Discover the Magic : Delia Owens "Where the Crawdads Sing" This evocative coming-of-age story follows Kya Clark, a young woman who grows up alone in the marshes of North Carolina. Owens spins a tale of resilience, survival, and the transformative power of nature, captivating readers with its evocative prose and mesmerizing setting. These popular novels represent just a fraction of the literary treasures that have emerged in 2023. Whether you seek tales of romance, adventure, or personal growth, the world of literature offers an abundance of captivating stories waiting to be discovered. The novel begins with Richard Papan, a bright but troubled young man, arriving at Hampden College. Richard is immediately drawn to the group of students who call themselves the Classics Club. The club is led by Henry Winter, a brilliant and charismatic young man. Henry is obsessed with Greek mythology and philosophy, and he quickly draws Richard into his world. The other members of the Classics Club are equally as fascinating. Bunny Corcoran is a wealthy and spoiled young man who is always looking for a good time. Charles Tavis is a quiet and reserved young man who is deeply in love with Henry. Camilla Macaulay is a beautiful and intelligent young woman who is drawn to the power and danger of the Classics Club. The students are all deeply in love with Morrow, and they are willing to do anything to please him. Morrow is a complex and mysterious figure, and he seems to be manipulating the students for his own purposes. As the students become more involved with Morrow, they begin to commit increasingly dangerous acts. The Secret History is a exceptional and thrilling novel that will keep you wondering until the very end. The novel is a cautionary tale about the dangers of obsession and the power of evil.

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